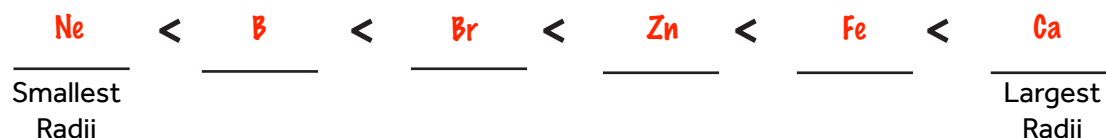


ATOM FORMATION

Atomic Number	20	10	30	5	26	35
Proton #	20	10	30	5	26	35
Electron #	20	10	30	5	26	35
# of energy levels	4	2	4	2	4	4
# of valence electrons	2	8	2	3	2	7
Electron configuration	1s ² 2s ² 2p ⁶ 3s ² 3p ⁶ 4s ²	1s ² 2s ² 2p ⁶	1s ² 2s ² 2p ⁶ 3s ² 3p ⁶ 4s ² 3d ¹⁰	1s ² 2s ² 2p ¹	1s ² 2s ² 2p ⁶ 3s ² 3p ⁶ 4s ² 3d ⁶	1s ² 2s ² 2p ⁶ 3s ² 3p ⁶ 4s ² 3d ¹⁰ 4p ⁵
Atomic symbol	Ca	Ne	Zn	B	Fe	Br
Element name	Calcium	Neon	Zinc	Boron	Iron	Bromine

Order the atoms above from smallest to largest atomic radii by placing atomic symbols on the lines below.



Summary Questions:

- Using a periodic table and the information collected above, describe the trend in atomic size across a period.

Atomic size decreases across a period.

- Using a periodic table and the information collected above, describe the trend in atomic size down a group.

Atomic size increases down a group.